

PCT

WORLD INTELLECTUAL PROPERTY ORGANIZATION
International Bureau



INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification ⁶ : A61K 33/24	A1	(11) International Publication Number: WO 96/30029 (43) International Publication Date: 3 October 1996 (03.10.96)
(21) International Application Number: PCT/GB96/00575 (22) International Filing Date: 19 March 1996 (19.03.96) (30) Priority Data: 9506126.3 25 March 1995 (25.03.95) GB (71) Applicant (for all designated States except US): JOHNSON MATTHEY PUBLIC LIMITED COMPANY [GB/GB]; 78 Hatton Garden, London EC1N 8JP (GB). (72) Inventors; and (75) Inventors/Applicants (for US only): MURRER, Barry, Anthony [GB/GB]; 17 Carnarvon Road, Reading, Berkshire RG1 5SB (GB). POWELL, Nigel, Anthony [GB/GB]; 4 Ibstock Close, Reading, Berkshire RG3 2NU (GB). (74) Agents: BREWER, Leonard, Stuart et al.; Johnson Matthey plc, Technology Centre, Blounts Court, Sonning Common, Reading RG4 9NH (GB).		(81) Designated States: AL, AM, AT, AU, AZ, BB, BG, BR, BY, CA, CH, CN, CZ, DE, DK, EE, ES, FI, GB, GE, HU, IS, JP, KE, KG, KP, KR, KZ, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, TJ, TM, TR, TT, UA, UG, US, UZ, VN, ARIPO patent (KE, LS, MW, SD, SZ, UG), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG). Published <i>With international search report. Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.</i>
(54) Title: PHARMACEUTICAL COMPOSITION CONTAINING SELECTED LANTHANUM CARBONATE HYDRATES (57) Abstract Selected lanthanum carbonate hydrates may be administered into the gastrointestinal tract, to treat hyperphosphataemia in patients with renal failure.		

FOR THE PURPOSES OF INFORMATION ONLY

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AM	Armenia	GB	United Kingdom	MW	Malawi
AT	Austria	GE	Georgia	MX	Mexico
AU	Australia	GN	Guinea	NE	Niger
BB	Barbados	GR	Greece	NL	Netherlands
BE	Belgium	HU	Hungary	NO	Norway
BF	Burkina Faso	IE	Ireland	NZ	New Zealand
BG	Bulgaria	IT	Italy	PL	Poland
BJ	Benin	JP	Japan	PT	Portugal
BR	Brazil	KE	Kenya	RO	Romania
BY	Belarus	KG	Kyrgyzstan	RU	Russian Federation
CA	Canada	KP	Democratic People's Republic of Korea	SD	Sudan
CF	Central African Republic	KR	Republic of Korea	SE	Sweden
CG	Congo	KZ	Kazakhstan	SG	Singapore
CH	Switzerland	LI	Liechtenstein	SI	Slovenia
CI	Côte d'Ivoire	LK	Sri Lanka	SK	Slovakia
CM	Cameroon	LR	Liberia	SN	Senegal
CN	China	LT	Lithuania	SZ	Swaziland
CS	Czechoslovakia	LU	Luxembourg	TD	Chad
CZ	Czech Republic	LV	Latvia	TG	Togo
DE	Germany	MC	Monaco	TJ	Tajikistan
DK	Denmark	MD	Republic of Moldova	TT	Trinidad and Tobago
EE	Estonia	MG	Madagascar	UA	Ukraine
ES	Spain	ML	Mali	UG	Uganda
FI	Finland	MN	Mongolia	US	United States of America
FR	France	MR	Mauritania	UZ	Uzbekistan
GA	Gabon			VN	Viet Nam

PHARMACEUTICAL COMPOSITION CONTAINING SELECTED LANTHANUM CARBONATE HYDRATES

This invention concerns a novel and inventive pharmaceutical composition and method, more particularly it concerns a composition for the treatment of hyperphosphataemia.

Hyperphosphataemia is a particular problem of patients with renal failure, using dialysis equipment. Conventional dialysis fails to reduce levels of phosphate in the blood, so that the levels rise in time. It is known to control phosphate levels by the oral administration of aluminium salts, or calcium salts. With the known toxic effects of aluminium, aluminium-based therapy tends to be avoided. In the case

of calcium salts, calcium is absorbed rather readily from the gut, and in turn causes hypercalcaemia.

It has been suggested (Nakagawa *et al*, Trans Am Soc Intern Organs, 31, (1985) 155-9) that hydrous cerium oxide could be used as a bead in an ion-exchange column, to bind phosphate during dialysis. Japanese published patent application 61 004 529 appears to cover the same idea, suggesting that the hydrous oxides of La, Ce and Y may be used in the column. However, although the rare earths are generally considered of low toxicity according to the Hodge-Sternier classification system (Am Ind Hyg Assoc Quart, 10, (1943), 93), their toxicity when given *iv*, which corresponds to use in a blood dialysis system, is significant and we are not aware that the suggested ion exchange system or any development thereof has met with widespread acceptance or has been tested clinically for hyperphosphataemia.

It appears that cerium oxide or oxalate was administered many years ago for different medical indications, but that this has fallen into complete disuse.

Japanese published patent application number 62-145024 (Asahi Chemical Ind KK) discloses that rare earth carbonates, bicarbonates or organic acid compounds may be used as phosphate binding agents. One example of said published application relates to the use of lanthanum carbonate, although in the tests described, cerium organic acid salts and carbonate gave better phosphate ion extraction than lanthanum carbonate. Example 11 of said published application prepares

$\text{La}_2(\text{CO}_3)_3 \cdot \text{H}_2\text{O}$, ie the monohydrate; all the other Examples are directed to rare earth carbonates other than lanthanum carbonate.

We have now discovered that certain forms of lanthanum carbonate exhibit improved performance in a variety of tests, over standard commercial lanthanum carbonate, which is believed to be the octahydrate form, and over $\text{La}_2(\text{CO}_3)_3 \cdot \text{H}_2\text{O}$ or similar compounds.

According to one aspect therefore, the present invention is the use of lanthanum carbonate of formula $\text{La}_2(\text{CO}_3)_3 \cdot x\text{H}_2\text{O}$ where x has a value from 3 to 6, preferably from 3.5 to 5, more especially from 3.8 to 4.5, for the preparation of a medicament for the treatment of hyperphosphataemia by administration into the gastrointestinal tract.

The invention further provides a pharmaceutical composition comprising said lanthanum carbonate, in admixture or association with a pharmaceutically acceptable diluent or carrier, in a form for administration into the gastrointestinal tract for the treatment of hyperphosphataemia.

The invention may also be expressed as a method of treatment of hyperphosphataemia in a patient with renal failure, comprising the administration of an effective dose of said lanthanum carbonate into the gastrointestinal tract.

According to another aspect, the present invention is a process for the preparation of lanthanum carbonate which comprises the steps of:

- (i) reacting lanthanum oxide with an acid which gives a soluble salt of lanthanum;
- (ii) reacting a solution of the thus obtained lanthanum salt with an alkali metal carbonate to produce a wet cake of lanthanum carbonate octahydrate; and
- (iii) controlled drying of the wet cake of lanthanum carbonate octahydrate so as to obtain a lanthanum carbonate with 3 to 6 molecules of water of crystallisation.

According to yet another aspect, the present invention is lanthanum carbonate when obtained by the above-mentioned process.

According to a further aspect, the present invention is lanthanum carbonate of the formula $\text{La}_2(\text{CO}_3)_3 \cdot x\text{H}_2\text{O}$ where x has a value from 3 to 6.

Embodiments of the present invention are described below, by way of example only, with reference to the accompanying drawings in which:

Figure 1 illustrates the phosphate-binding capability of lanthanum carbonates having different degrees of water of crystallisation;

Figure 2 illustrates the drying curves for five batches of lanthanum carbonate prepared by the method indicated in Example 1;

Figure 3 illustrates the XRD analysis of lanthanum carbonate $4\text{H}_2\text{O}$ prepared by the method indicated in Example 2; and

Figure 4 illustrates the XRD analysis of lanthanum carbonate $8.8\text{H}_2\text{O}$ of Sample 1 above.

5

For the tests described hereinafter, samples of lanthanum carbonate were obtained as follows:

Sample 1. Commercial lanthanum carbonate obtained from a chemical company.

10 This was characterised by elemental analysis (La, C, H), TGA, X-ray powder diffraction and ir spectroscopy, to have the formula $\text{La}_2(\text{CO}_3)_3 \cdot 8.8\text{H}_2\text{O}$.

Samples 2 - 4 were prepared by heating portions of Sample 1 at varying temperatures for varying lengths of time, either under vacuum or at atmospheric pressure to obtain materials of formula $\text{La}_2(\text{CO}_3)_3 \cdot x\text{H}_2\text{O}$ where $0 < x < 8$.

15

Sample	Initial wt (g)	Temp (°C)	Time (min)	Vacuum (Y/N)	Wt loss (g)	x
2	5.00	175	240	Y	1.09	1.3
3	20.0	80	180	N	2.6	4.4
4	5.01	100	720*	N	0.96	2.2

20

* Dried to constant weight.

Sample 5 is a sample of lanthanum carbonate which when analysed indicated a formula of $\text{La}_2(\text{CO}_3)_3 \cdot 4\text{H}_2\text{O}$.

Sample 6 is a sample of lanthanum carbonate prepared according to Example 1 below and having the formula $\text{La}_2(\text{CO}_3)_3 \cdot 3.8\text{H}_2\text{O}$.

In order to show that certain lanthanum carbonate hydrates are significantly different in phosphate binding activity from both lanthanum carbonate octahydrate and from $\text{La}_2(\text{CO}_3)_3 \cdot \text{H}_2\text{O}$, samples were tested as follows:

10

i) a stock solution was prepared by dissolving 13.75g of anhydrous Na_2HPO_4 , 8.5g of NaCl in 1 litre deionised water.

ii) 100ml of the stock solution was adjusted to pH3 by the addition of concentrated HCl.

15

iii) A 5ml sample was taken and filtered through a $0.02\mu\text{m}$ filter to give a Time 0 sample. This was analysed for phosphate using a Sigma Diagnostics Colorimetric Phosphorus test kit.

iv) 5ml fresh stock solution was added to re-establish 100ml, and the pH was re-adjusted to approximately 3.

20

v) $\text{La}_2(\text{CO}_3)_3 \cdot x\text{H}_2\text{O}$ as a dry powder was added in an amount according to the molecular weight of the particular hydrate, to give a two-fold molar excess of lanthanum over phosphate and stirred at room temperature.

vi) Sampling was carried out at time intervals from 0.5 to 10 minutes, and the percentage of phosphate was determined as in iii) above. The results are shown in the Table 1 below.

TABLE 1

TIME (Minutes)	% PHOSPHATE REMOVED					
	Sample					
	1	2	3	4	5	6
0						
0.5		13.4	18.8	15.1	22.9	31.4
1	29	18.4	31.5	26.8	40.4	55.5
1.5		25.4	43.1	36	55.2	74.8
2		28.1	50.6	45.3	69.5	88.1
2.5		30.8	60.5	51.8	79.9	95.3
3		34.4	69	57.6	90.3	99.6
4						100
5	70.5	39.9	96.5	76.3	100	100
10	100	ND	99.1	ND	100	100

It can readily be seen from Table 1 that Sample 3 ($\text{La}_2(\text{CO}_3)_3 \cdot 4.4\text{H}_2\text{O}$); Sample 5 ($\text{La}_2(\text{CO}_3)_3 \cdot 4\text{H}_2\text{O}$) and Sample 6 ($\text{La}_2(\text{CO}_3)_3 \cdot 3.8\text{H}_2\text{O}$) bind phosphate appreciably quicker than the $8.8\text{H}_2\text{O}$, $1.3\text{H}_2\text{O}$ or $2.2\text{H}_2\text{O}$ forms. We believe that the results for $\text{La}_2(\text{CO}_3)_3 \cdot 1.3\text{H}_2\text{O}$ are in agreement with the results shown in the above mentioned Japanese published patent application number 62-145024 where for $\text{La}_2(\text{CO}_3)_3 \cdot \text{H}_2\text{O}$, only 90% removal is shown after 120 minutes.

It can also be readily seen from Figure 1 of the accompanying drawings that the highest phosphate removal is obtained with lanthanum carbonates having 3 to 6 molecules of water.

5 The present invention offers the possibility of binding phosphate without any incursion of lanthanum into the blood stream, where toxic effects can cause problems. The specified lanthanum carbonate has negligible absorption from the gut, as shown by the *in vivo* tests described below.

10 Throughout this document, the term "treatment" is intended to include preventative treatment.

Processes for preparing lanthanum carbonates according to the present invention are described by way of illustration in the following Examples 1 and 2.

15

EXAMPLE 1

20 Lanthanum oxide (1.5kg, 4.58mol) was suspended in water (5.5 litres) in a 20 litre flask. Nitric acid (Analar grade, 69%, SG 1.42, 1.88 litres, 29.23mol) was added to the stirred solution over 1.5 hours at such a rate as to keep the temperature between 60-80°C. The resulting lanthanum nitrate solution was left to cool to room temperature and filtered. A solution of sodium carbonate (1.65kg, 15.57mol) in water (7.75 litres) was added to the stirred lanthanum nitrate solution over 45 minutes. At the end of the addition the pH of the suspension was 9.74. The suspension was left

overnight, filtered (Buchner funnel, 540 paper) and dried on the filter in a current of air for 30 minutes. The solid was then re-suspended in water, stirred for 40 minutes and filtered. This procedure was repeated to give a total of six washes, when the nitrate concentration in the filtrate was <500ppm. The final material (4.604kg) was divided between three Pyrex dishes and a sample from each analysed for water content. (By decomposition of weighed sample of $(\text{La}_2(\text{CO}_3)_3 \cdot x\text{H}_2\text{O})$ at 1050°C , 2 hours to La_2O_3). The dishes were then placed in a fan oven at 80°C and the weight loss of each dish monitored until the material of the required degree hydration was obtained. The progress of the drying is shown below

Time (hours)	mol $\text{H}_2\text{O}/\text{La}$		
	Dish 1	Dish 2	Dish 3
3.50	10.9	13.5	12.6
12	5.7	6.0	5.2
14	5.3	5.4	4.6
16	4.9	5.1	4.3
17	4.4	4.6	3.8
19.5	3.8	4.0	3.2

Drying curves for five batches produced by this route are shown in

Figure 2.

$\text{La}_2(\text{CO}_3)_3 \cdot 3.8\text{H}_2\text{O}$ from dish 1 was selected as Sample 6 for the phosphate binding tests set forth in Table 1.

EXAMPLE 2

The process of Example 1 was repeated but using hydrochloric acid (12.28M, 2.48 litres) in place of nitric acid to dissolve lanthanum oxide (1.5kg). The yield of crude product after six washes was 4.378kg. The product was divided in three approximately equal portions in Pyrex dishes and dried in a fan oven at 80°C. After 2 hours a sample was taken from each tray and water analysed by decomposition to lanthanum oxide as described above. These figures were used to calculate the weight loss needed to give material of the required composition. The time course of the drying process is shown below.

Time (hours)	mol H ₂ O/La		
	Dish 1	Dish 2	Dish 3
2	21.3	22.1	20.4
5.5	12.3	13.2	12.2
9	7.9	8.0	7.6
11.5	6.9	7.0	6.6
17	4.9	5.1	4.6
18.5	4.6	4.8	4.2
19.5	4.4	4.6	4.1
20	4.3	4.6	4.0

Samples were taken from each dish, combined and analysed. The following results were obtained:

	Found	Calculation for $\text{La}_2(\text{CO}_3)_3 \cdot 4\text{H}_2\text{O}$
% La (gravimetric)	52.38%	52.4%
carbonate (titration)	5.76mol/g	5.66mol/g
H_2O (NMR)	13.06%	13.59%

5

The XRD analysis for lanthanum carbonate $4\text{H}_2\text{O}$ prepared by the method of Example 2 is illustrated in Figure 3.

10

Figure 4 illustrates the XRD of lanthanum carbonate $8.8\text{H}_2\text{O}$ and it is evident that it has a different crystalline structure from lanthanum carbonate $4\text{H}_2\text{O}$ prepared by the method of Example 2. The XRD analysis of lanthanum carbonate $4\text{H}_2\text{O}$ prepared by the method of Example 1 was similar to the XRD analysis of lanthanum carbonate $4\text{H}_2\text{O}$ prepared by the method of Example 2.

15

Pharmaceutical compositions for oral administration according to the invention may be formulated and manufactured using methods well known in the art. Suitable diluents or carriers are also well known. The compositions may desirably be in a dosage form, to provide a single daily dose, or a number of sub-daily dosages. Conventional pharmacological methods may be used to ascertain suitable dose levels.

20 The level of phosphate in the food that an individual ingests is important. Daily dosages are indicated to be in the range 0.1 to 50g, preferably about 0.5 to 15g. Suitable forms for oral administration include solid forms such as tablets, capsules and dragees and liquid forms such as suspensions or syrups. In addition to diluents and carriers, it is conventional in the formulation of oral preparations to include non-active

ingredients such as thickeners, taste-improving components and colouring agents. The said carbonate may also be coated or treated to provide delayed-release forms. Preferably, the required daily dosage is given in tablet form, eg chewable tablet form, to be taken with meals. A suitable daily dosage of about 2g for 70kg man, should be compared with a daily dosage of 20g for a commercial calcium-based phosphate binding composition.

To demonstrate that the lanthanum carbonate of the invention (or lanthanum phosphate formed after binding to phosphate in the gut) is fully excreted and does not pass out of the gut into the circulation system when given orally, three rats were dosed with 20mg/kg of $\text{La}_2(\text{CO}_3)_3 \cdot 4\text{H}_2\text{O}$ (Sample 5) and kept in metabolic cages where faeces and urine could be collected. The results are shown in Table 2 below.

Animal No.	Time (hours)	% La Recovered
1	24	103.2
1	48	0.1
1	72	<0.2
1	Total	103.3
2	24	75.3
2	48	23
2	72	1.2
2	Total	99.5

Animal No	Time (hours)	%La Recovered
3	24	93.8
3	48	10
3	72	0.1
3	Total	103.8

5

It can be seen that after 72 hours, all of the lanthanum has been excreted. In the urine samples, the amount of lanthanum was below detection limits.

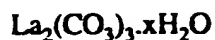
10

After the test, the rats were sacrificed, and kidney, liver and femur were analysed for lanthanum. In all cases, the amount of lanthanum was below 0.1ppm.

CLAIMS

1. A pharmaceutical composition for the treatment of hyperphosphataemia, comprising lanthanum carbonate of formula

5



where x has a value from 3 to 6, in admixture or association with a pharmaceutically acceptable diluent or carrier.

10

2. A composition according to claim 1, wherein in the lanthanum carbonate, x has a value from 3.5 to 5.

3. A composition according to claim 2, wherein in the lanthanum carbonate, x has a value from 3.8 to 4.5.

15

4. A composition according to any one of claims 1 to 3, in a form suitable for oral administration.

5. A composition according to any one of claims 1 to 4 in unit dosage form to provide from 0.1 to 20g/day.

20

6. The use of lanthanum carbonate as defined in any one of claims 1 to 3, for the preparation of a medicament for the treatment of hyperphosphataemia by administration into the gastrointestinal tract.

7. A process for the preparation of lanthanum carbonate as defined in any one of claims 1 to 3 which comprises the steps of:

(i) reacting lanthanum oxide with an acid which gives a soluble salt of lanthanum;

(ii) reacting a solution of the thus obtained lanthanum salt with an alkali metal carbonate to produce a wet cake of lanthanum carbonate octahydrate; and

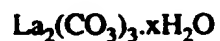
(iii) controlled drying of the wet cake of lanthanum carbonate octahydrate so as to obtain a lanthanum carbonate with 3 to 6 molecules of water of crystallisation.

8. A process as claimed in claim 7 wherein the acid is nitric acid or hydrochloric acid.

9. A process as claimed in claim 7 or 8 wherein the alkali metal carbonate is sodium carbonate.

10. Lanthanum carbonate prepared according to the process of any of claims 7, 8 or 9.

11. Lanthanum carbonate of the formula



where x has a value from 3 to 6.

Fig. 1

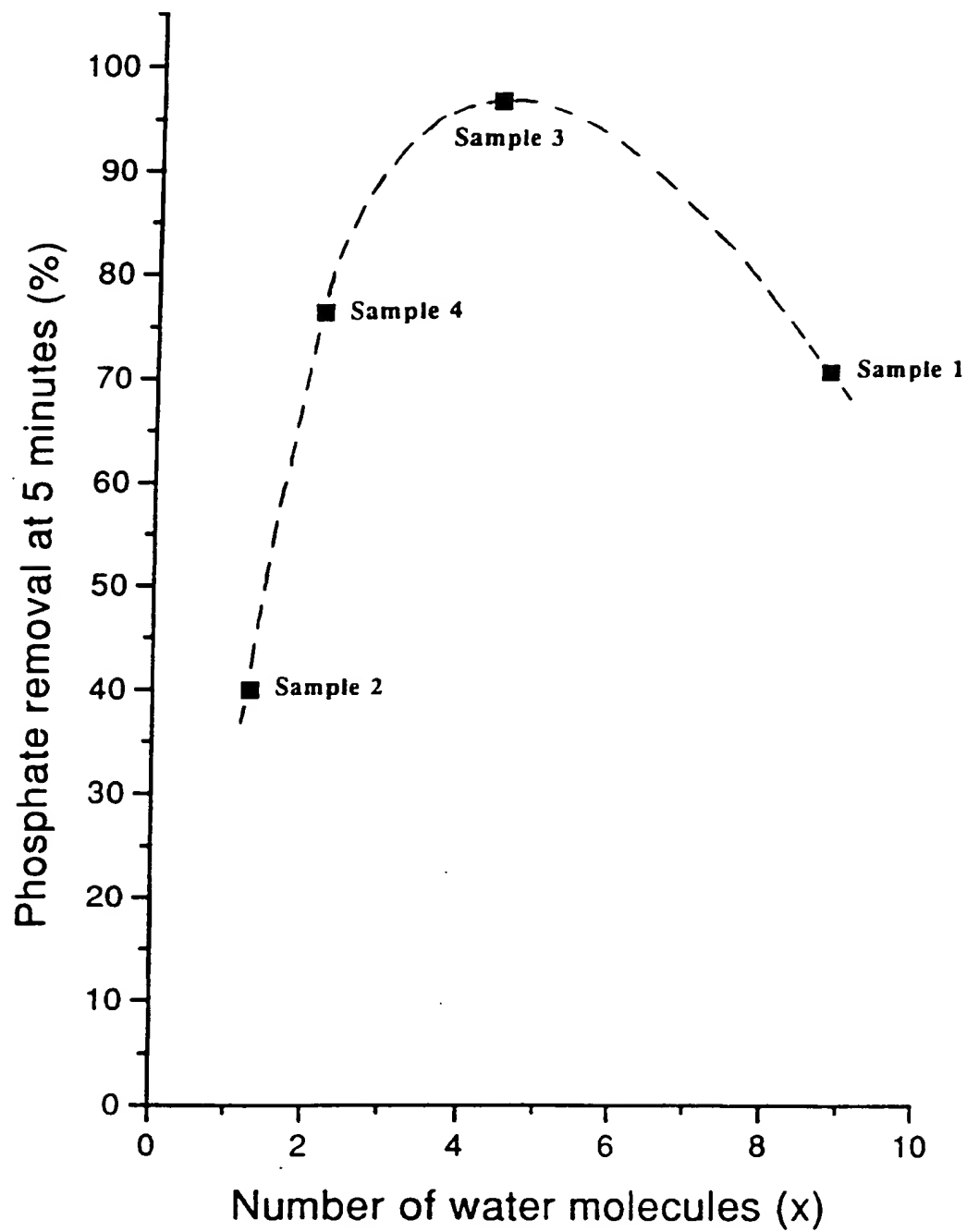


Fig. 2

2/4

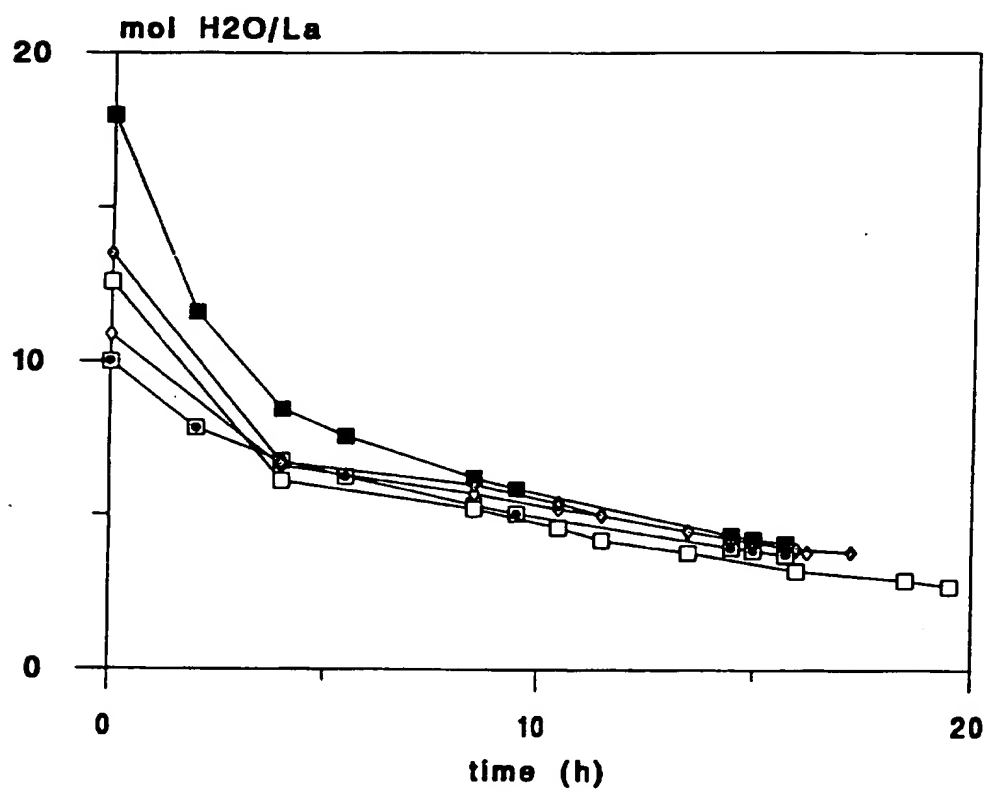


Fig. 3

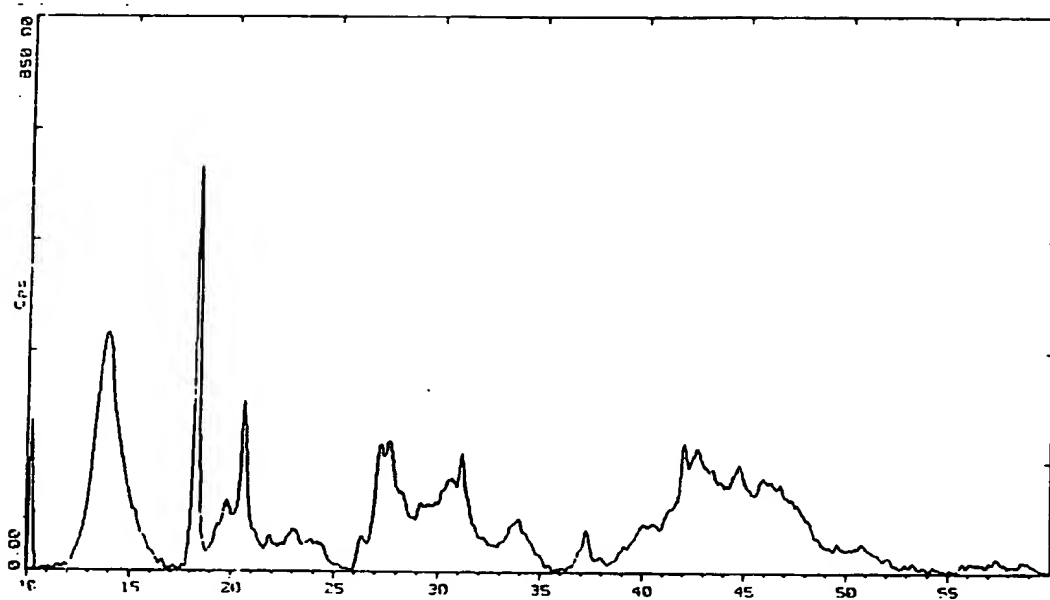
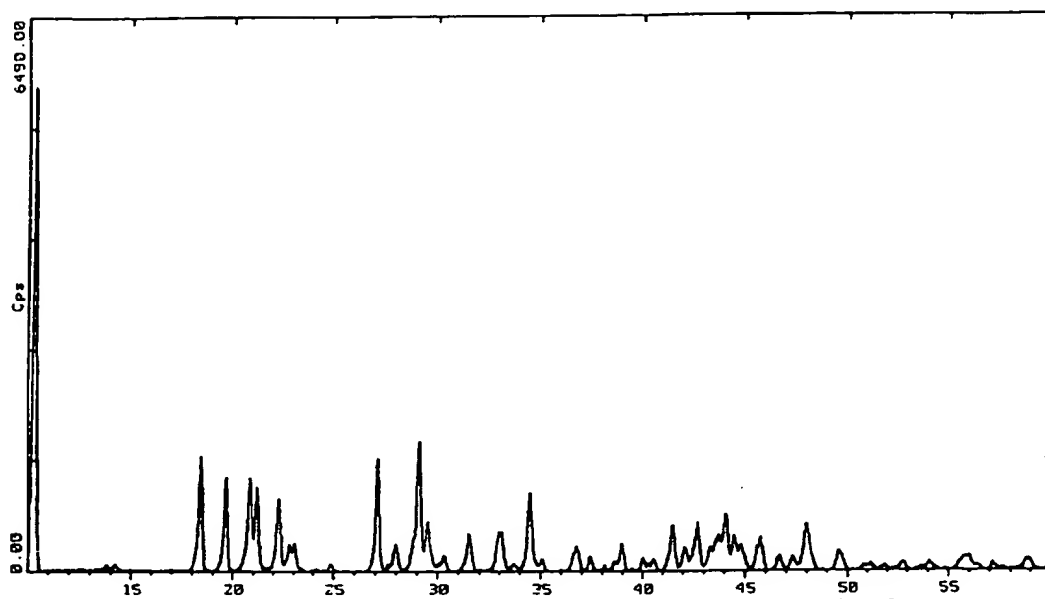


Fig.4



INTERNATIONAL SEARCH REPORT

national Application No
PCT/GB 96/00575A. CLASSIFICATION OF SUBJECT MATTER
IPC 6 A61K33/24

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
IPC 6 C01F A61K

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	PATENT ABSTRACTS OF JAPAN vol. 011, no. 371 (C-462), 3 December 1987 & JP,A,62 145024 (ASAHI CHEM IND CO LTD), 29 June 1987, cited in the application see abstract	1-11
X	--- JOURNAL OF THE LESS-COMMON METALS, vol. 167, no. 2, 1 January 1991, pages 223-232, XP000202645 NAOHISA YANAGIHARA ET AL: "SYNTHESIS OF LANTHANIDE CARBONATES" see page 226; table 2 --- -/--	11

☒ Further documents are listed in the continuation of box C.☐ Patent family members are listed in annex.

* Special categories of cited documents:

- *A* document defining the general state of the art which is not considered to be of particular relevance
- *E* earlier document but published on or after the international filing date
- *L* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- *O* document referring to an oral disclosure, use, exhibition or other means
- *P* document published prior to the international filing date but later than the priority date claimed

T later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

X document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

Y document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art

A document member of the same patent family

Date of the actual completion of the international search

13 August 1996

Date of mailing of the international search report

29.08.96

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2
NL - 2280 HV Rijswijk
Tel. (+ 31-70) 340-2040, Tx. 31 651 epo nl,
Fax (+ 31-70) 340-3016

Authorized officer

Leherte, C

INTERNATIONAL SEARCH REPORT

International Publication No

PCT/GB 96/00575

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	CHEMICAL ABSTRACTS, vol. 107, no. 26, 28 December 1987 Columbus, Ohio, US; abstract no. 249009, MINEELY, PATRICK J.; TARIQ, SHABBIR A.: "Molten potassium pyrosulfate: reactions of lanthanum metal and six of its compounds" XP002010788 see abstract & AUST. J. CHEM. (1987), 40(7), 1309-14, ---	11
X	CHEMICAL ABSTRACTS, vol. 104, no. 26, 30 June 1986 Columbus, Ohio, US; abstract no. 236218, MZAREULISHVILI, N. V.; NATIDZE, V. P.: "Study of interaction of lanthanum nitrate with alkali metal and ammonium carbonates" XP002010789 see abstract & SOOBASHCH. AKAD. NAUK GRUZ. SSR (1986), 121(1), 81-4, ---	11
X	CHEMICAL ABSTRACTS, vol. 87, no. 20, 14 November 1977 Columbus, Ohio, US; abstract no. 161013, ODA, TOSHIYUKI: "Studies on the crystal water of lanthanum carbonates" XP002010790 see abstract & OITA DAIGAKU KYOIKUGAKUBU KENKYU KIYO, SHIZEN KAGAKU (1975), 4(5), 1-6, ---	11
X	CHEMICAL ABSTRACTS, vol. 87, no. 18, 31 October 1977 Columbus, Ohio, US; abstract no. 142122, KARAPET'YANTS, M. KH.; MAIER, A. I.; BAS'KOVA, N. A.: "Standard heats of formation of rare earth element and yttrium carbonates" XP002010791 see abstract & IZV. AKAD. NAUK SSSR, NEORG. MATER. (1977), 13(7), 1279-81, -----	11